

A2 Music



REVISION GUIDE

Perfect preparation for exams from 2012

Alistair Wightman

2nd edition

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Alistair Wightman, Roy Wightman



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now part of OCR and Edexcel for whom he pioneered the introduction of Music Technology as an A level subject **Edexcel A2 Music Revision Guide** Jonny Martin, 2011-02-11 This Edexcel A2 Music Revision Guide is designed to prepare you fully for Sections B and C of your final A2 Music exam For exams 2012 onwards Written in a clear and concise manner by an experienced examiner and teacher it includes an introduction to the exam that explains what you ll be expected to know a succinct revision section that presents all the key facts for each set work broken down into individual segments on the elements of music as well as sample essay questions mark schemes and answers This guide will help you to understand how the exam works how questions are worded and what your examiners are looking for With helpful hints on how to improve your written answers and a comprehensive glossary it will give you the edge you need to achieve a better grade **A2-Level Biology Edexcel Complete Revision & Practice** Richard Parsons, 2009-06 A2 Level Biology Edexcel Complete Revision Practice **Edexcel AS Chemistry Revision Guide** Phillip Dobson, David Craggs, Geoff Wright, 2009 Drawing on exam data from Edexcel s Results Plus service this book offers tailored support for the 2008 specification It includes exam style questions worked examples guidance on practical assessment case studies and investigations Thinking tasks quick questions and checklists help students track their progress Edexcel AS Biology Revision Guide Gary Skinner, Stephen Winrow-Campbell, John Dunkerton, 2009 Written by experienced examiners this revision guide for AS provides tailored support for the 2008 specification Featuring guidance from the people who write and mark exam papers and drawing on real exam data from Edexcel s ResultsPlus service students can be confident they are designed to specifically meet their needs

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and other countries of the British Commonwealth This qualification is designed for students aged 16 18 and is typically taken after completing the General Certificate of Secondary Education GCSE level of study The qualification assesses the student s proficiency in Chinese language skills including listening speaking reading and writing and is available in both spoken Mandarin and spoken Cantonese The Pearson Edexcel Level 3 Advanced GCE in Chinese spoken Mandarin spoken Cantonese consists of two externally examined papers assessing listening reading and writing and a speaking assessment The first paper 9CN0 01 evaluates listening reading and translation skills while the second paper 9CN0 02 tests the student s ability to write responses to works and translation The third paper 9CN0 03M 03C is an internally conducted and externally assessed speaking assessment This advanced level qualification is ideal for students who possess a strong command of the Chinese language and wish to demonstrate their proficiency to universities or employers Edexcel GCE A Level Chinese Paper 3 Speaking assessment is conducted internally by the school or institution but is externally assessed by a qualified examiner The total assessment time is between 27 and 30 minutes which includes a single period of 15 minutes formal preparation time The assessment is worth 30% of the overall qualification with a total of 72 marks available The assessment is divided into two tasks Task 1 requires students to read and respond to two short texts based on Theme 1 and then hold a discussion Task 2 is based on a discussion from Theme 2 which assesses the students ability to use a range of language accurately and effectively as well as their knowledge and understanding about the culture and society of the Chinese speaking world Overall the assessment is designed to assess a range of skills and abilities including language accuracy communication and interaction skills summarising and understanding written sources in speech and knowledge and understanding about the culture and society of the Chinese speaking world The assessment aims to provide a comprehensive evaluation of the students language proficiency and cultural awareness and to help them develop their skills further

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow 2\text{Cr}^{3+} \text{aq} + 7\text{H}_2\text{O} \text{l}$ s label 4 2 3

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+
 net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions
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web the six Na^+ ions and the six Cl^- ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{aq} + \text{KCl} \text{aq} \rightarrow \text{AgCl} \text{s} + \text{KNO}_3 \text{aq}$ total ionic $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$ net ionic $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s}$ 2

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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