

**18.3****COLLIGATIVE PROPERTIES OF SOLUTIONS****SECTION REVIEW****Objectives**

- Explain on a particle basis why a solution has a lower vapor pressure than the pure solvent of that solution.
- Explain on a particle basis why a solution has an elevated boiling point and a depressed freezing point compared with the pure solvent.

**Key Terms**

- colligative properties
- boiling-point elevation
- freezing-point depression

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or even line.

The effects in solution of a nonvolatile \_\_\_\_\_ on the properties of the solvent are called \_\_\_\_\_. They include \_\_\_\_\_ point and vapor pressure \_\_\_\_\_, and boiling-point \_\_\_\_\_. In each case, the magnitude of the effect is \_\_\_\_\_ proportional to the number of solute molecules or ions present in the \_\_\_\_\_.

Colligative properties are a function of the number of solute \_\_\_\_\_ in solution. For example, one mole of sodium chloride produces \_\_\_\_\_ as many particles in solution as one mole of sucrose and, thus, will depress the freezing point of water \_\_\_\_\_ as much.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
7. \_\_\_\_\_
8. \_\_\_\_\_
9. \_\_\_\_\_
10. \_\_\_\_\_

**Part B True-False**

Classify each of these statements as an always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 11. When added to 1000 g of water, 2 moles of a solute will decrease the boiling point by  $0.212^{\circ}\text{C}$ .
- \_\_\_\_\_ 12. One mole of solute A will depress the freezing point of 1000 g of water the same as one mole of solute B.

# Colligative Properties Of Solutions Section Review Key

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