A useful way to compare reaction rates is to record the time required for half of the reactants to react, called the half-life of the reaction. With the Temperature set to 200 *C, click Play. Click Pause (__) when the number of reactant molecules is 10. Record

the half-life time in the first space of the table below. Repeat the experiment at different temperatures to complete the table.

Trial	200 °C	150 °C	100 °C	50 °C
1	05:45	04:45	23:18	174:20
2	13:49	07:00	20:48	244:45

What do your results indicate?

The lower the temperature, the longer the time for the reactants to change.

2. How will increasing the reactant concentration affect the rate of the reaction? Why?

Increasing the concentration of one or more reactants will often increase the rate of reaction. This occurs because a higher concentration of a reactant will lead to more collisions of that reactant in a specific time period. 2. Physical state of the reactants and surface area.

Set the reactant concentration to 2.0 mol/L, keeping the temperature at 200 °C. Use
the Gizmo to measure the half-life of the reaction for each surface area setting. (There
will now be 20 reactant molecules left at the half-life.)

Trial	Minimum surface area	Maximum surface area
1	3:36	0:40
2	6:06	0:50

Why does the reaction proceed more quickly when the surface area is increased?

Collision Theory Lab Report

G Psacharopoulos

Collision Theory Lab Report:

The Enigmatic Realm of Collision Theory Lab Report: Unleashing the Language is Inner Magic

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