

COMBINED GAS LAW

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2} \quad \text{or} \quad P_1 V_1 T_2 = P_2 V_2 T_1$$

Complete the following chart.

| | P ₁ | V ₁ | T ₁ | P ₂ | V ₂ | T ₂ |
|---|----------------|----------------|----------------|----------------|----------------|-----------------|
| 1 | 1.5 atm | 3.0 L | 20° C | 2.5 atm | 1.2 L | 30° C |
| 2 | 720 torr | 256 mL | 25° C | 368.64 torr | 250 mL | 50° C |
| 3 | 600 mmHg | 2.5 L | 22° C | 760 mmHg | 1.8 L | 24.1228° C |
| 4 | 25000 atm | 750 mL | 0.0° C | 2.0 atm | 500 mL | 25° C |
| 5 | 95 kPa | 4.0 L | 315.76° C | 101 kPa | 6.0 L | 471 K or 198° C |
| 6 | 650. torr | 467.3 mL | 100° C | 900. torr | 225 mL | 150° C |
| 7 | 850 mmHg | 1.5 L | 15° C | 255 mmHg | 2.5 L | 30° C |
| 8 | 125 kPa | 125 mL | 48° C | 100 kPa | 100 mL | 75° C |

Combined Gas Law Answer Key With Work

G Orfield



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