#### Concentration - Understanding Molarity and Parameters affecting Molarity Remote Lab

(This lesson is designed for a student working remotely.)

This lab uses the <u>Concentration</u> and <u>Molarity</u> simulations from PhET interactive Simulations at University of Colorado Boulder, under the CC-BY 4.0 license.

https://phet.colorado.edu/sims/html/concentration/latest/concentration\_en.html https://phet.colorado.edu/sims/html/molarity/latest/molarity\_en.html

#### Learning Goals

- Describe the relationships between volume and amount of solute to solution concentration.
- Predict how solution concentration will change for any action (or combination of actions) that adds or removes water, solute, or solution, and explain why.

Develop your understanding: Explore the <u>Concentration</u> simulation. Try to find all the ways you can change the concentration of the solution in the beaker.



#### Explain your understanding:

- 1. Mole calculation review. Show your work:
  - a. How many moles is

55 g of NaCl? = 0.54 mol NaCl

0.025 g of NaCO<sub>3</sub>? = 0.0004 g of NaCO<sub>3</sub>

b. How many grams is

0.5 moles of NaCl? = 29.2 g

2.11 moles of NaCO<sub>2</sub>? = 123.3 g

- There are several ways to measure Concentration. This simulation uses Molarity (mol/L).
  - a. What does mol/L mean? mol/L means Molar Concentration.

b. Describe in your own words how mol/L compares to grams/liter. (similarities and differences)
 Similarities:

mol/L and grams/liter both are units of concentration.

#### Differences:

mol/L concentration says how many moles of solute are present in 1L of solution. grams/liter concentration says how many grams of solute are present in 1L of solution.

3. Describe at least 2 ways in the simulation to change each of the parameters:

# **Concentration And Molarity Lab**

Mr. Rohit Manglik

## **Concentration And Molarity Lab:**

Fundamental Laboratory Approaches for Biochemistry and Biotechnology Alexander J. Ninfa, David P. Ballou, Marilee Benore, 2009-05-26 Ninfa Ballou Benore is a solid biochemistry lab manual dedicated to developing research skills allowing students to learn techniques and develop the critical thinking and organizational approaches necessary to conduct laboratory research Ninfa Ballou Benore focuses on basic biochemistry laboratory techniques but also includes molecular biology exercises a reflection of most courses which concentrate on traditional biochemistry experiments and techniques The experiments are designed so that theory and technique are learned as fundamental research tools and the biochemistry and molecular biology applications are seamlessly integrated throughout the manual The manual also includes an introduction to ethics in the laboratory uncommon in similar manuals Most importantly perhaps is the authors three pronged approach to encouraging students to think like a research scientist first the authors introduce the scientific method and the hypothesis as a framework for developing conclusive experiments second the manual s experiments are designed to become increasingly complex in order to teach more advanced techniques and analysis finally gradually the students are required to devise their own protocols In this way students and instructors are able to break away from a cookbook approach and to think and investigate for themselves Suitable for lower level and upper level courses Ninfa spans these courses and can also be used for some first year graduate work Laboratory Manual for Biotechnology Verma, Ashish S./ Das Surajit & Singh Anchal, Laboratory Manual in Biotechnology Students **Mathematics for the Clinical Laboratory E-Book** Lorraine J. Doucette, 2020-06-18 Selected for Doody's Core Titles 2024 with Essential Purchase designation in Laboratory Technology Master the skills you ll need to perform accurate clinical laboratory calculations Mathematics for the Clinical Laboratory 4th Edition demonstrates the calculations used in the analysis of test specimens It begins by explaining basic mathematical principles and then covers the types of calculations needed in specific areas of the clinical lab including urinalysis hematology and microbiology Finally it focuses on the statistical calculations used in quality assurance and quality control Step by step examples reinforce your understanding and calculation templates and practice problems ensure that you make correct calculations every time Step by step examples explain basic mathematical principles and show you exactly how to perform each type of calculation Sample problems with answers can also be used as templates for solving laboratory calculations Practice problems at the end of each chapter provide a self assessment tool helping you determine what you need to review Summaries of important formulas are included at the end of the text s major sections Coverage of statistical calculations includes standard deviation as well as calculations associated with quality assurance and quality control Quick tips and notes make it easier to understand and remember pertinent information Learning objectives at the beginning of each chapter provide measurable outcomes to achieve by completing the chapter material Full color design includes 100 illustrations Useful appendix of Greek symbols provides a quick reference to turn to when studying Glossary at the back of

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