

Concentration and Molarity (M) table

Name: _____

Mass concentration (mg/L)

Using the concentration and the formula for Molarity in the table, complete the table below

Mass of compound	Volume (L)	Molarity of compound (M)	Mass of compound	Volume (L)	Molarity of compound (M)
100	10	0.007	0.005	10	0.0005
200	20	0.01	20	0.02	0.001
0.5	10	0.0005	0.5	0.05	0.0001
20	20	0.001	500.0	0.05	0.01

Concentration (Moles/L) and Molarity (Moles/L) and Molarity (Moles/L) table

1. What is the molarity of a solution made by adding 100 mg of NaCl to 10 L of water? **0.0007 M**
2. What is the molarity of a solution made by adding 200 mg of NaCl to 20 L of water? **0.0001 M**
3. What is the molarity of a solution made by adding 0.5 g of NaCl to 10 L of water? **0.000005 M**
4. What is the molarity of a solution made by adding 20 g of NaCl to 20 L of water? **0.0001 M**
5. What is the molarity of a solution made by adding 500.0 g of NaCl to 0.05 L of water? **0.000001 M**
6. What is the molarity of a solution made by adding 100 mg of NaCl to 10 L of water? **0.000001 M**
7. What is the molarity of a solution made by adding 200 mg of NaCl to 20 L of water? **0.000001 M**
8. What is the molarity of a solution made by adding 0.5 g of NaCl to 10 L of water? **0.000001 M**
9. What is the molarity of a solution made by adding 20 g of NaCl to 20 L of water? **0.000001 M**
10. What is the molarity of a solution made by adding 500.0 g of NaCl to 0.05 L of water? **0.000001 M**
11. What is the molarity of a solution made by adding 100 mg of NaCl to 10 L of water? **0.000001 M**
12. What is the molarity of a solution made by adding 200 mg of NaCl to 20 L of water? **0.000001 M**
13. What is the molarity of a solution made by adding 0.5 g of NaCl to 10 L of water? **0.000001 M**
14. What is the molarity of a solution made by adding 20 g of NaCl to 20 L of water? **0.000001 M**
15. What is the molarity of a solution made by adding 500.0 g of NaCl to 0.05 L of water? **0.000001 M**

Concentration And Molarity Phet Chemistry Labs

Answer Key

SA Adler



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