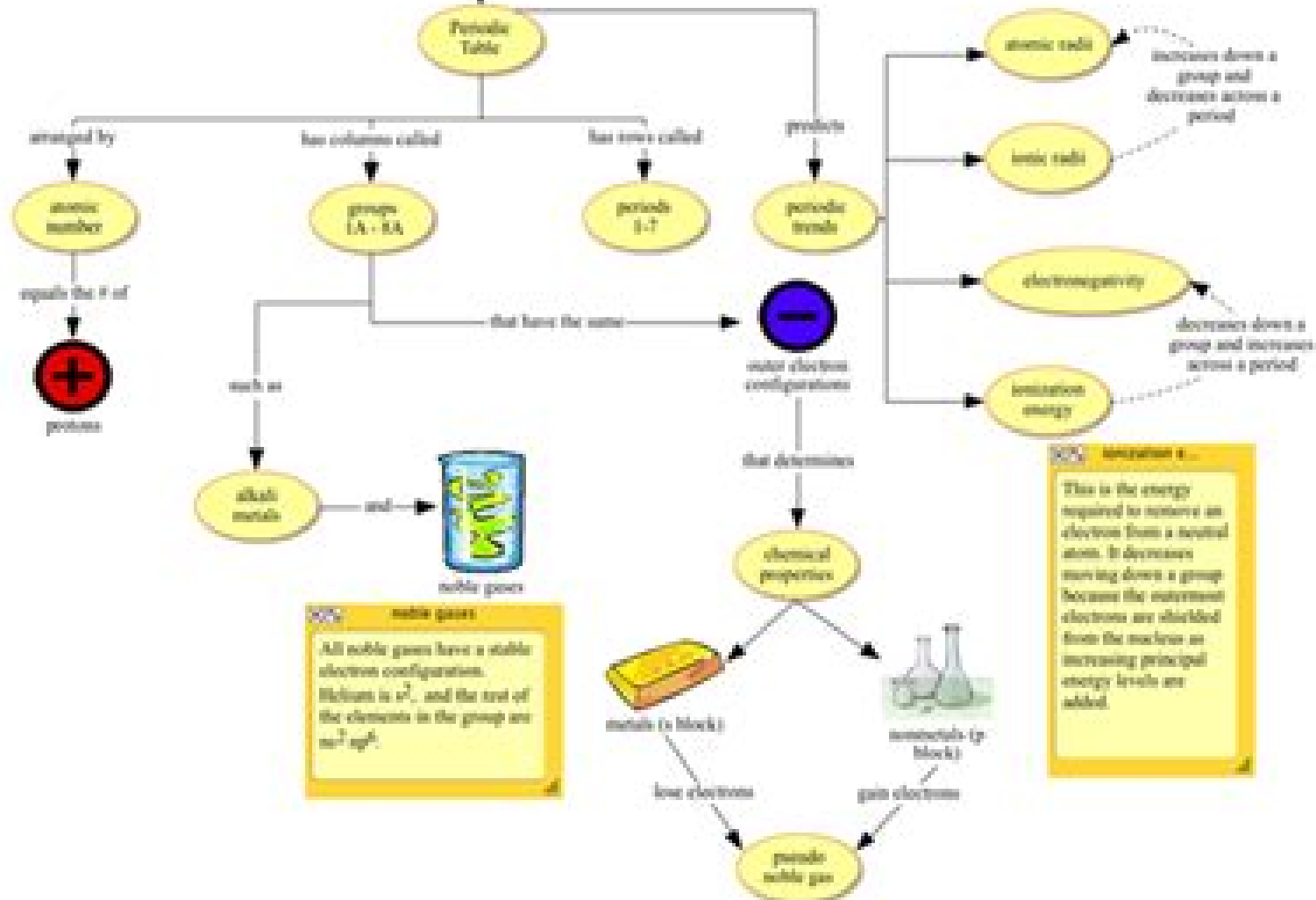




**noble gases**

All noble gases have a stable electron configuration. Helium is  $1s^2$ , and the rest of the elements in the group are  $ns^2 np^6$ .

**ionization energy**

This is the energy required to remove an electron from a neutral atom. It decreases moving down a group because the outermost electrons are shielded from the nucleus as increasing principal energy levels are added.

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**Daniel L. Hartl**



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