

Molarity:

- a quantitative description of solution concentration.
- Abbreviated M

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

$$4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}} \quad \text{moles of solute} = \boxed{48.0 \text{ mol}}$$

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

$$\begin{aligned} ? \text{ L} &= 250 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.25 \text{ L} \\ 0.150 \text{ M} &= \frac{\text{moles of solute}}{0.25 \text{ L}} \quad \text{moles of solute} = \boxed{0.038 \text{ mol}} \end{aligned}$$

3. What is the molarity of a solution of HNO_3 that contains 12.6 grams HNO_3 in 1.0 L of solution?

$$\begin{aligned} ? \text{ mol HNO}_3 &= 12.6 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} = 0.200 \text{ mol HNO}_3 \\ \text{M} &= \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = \boxed{0.200 \text{ M}} \end{aligned}$$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

$$\begin{aligned} 0.700 \text{ M} &= \frac{\text{moles of solute}}{0.250 \text{ L}} \quad \text{moles of solute} = 0.175 \text{ mol} \\ ? \text{ g KNO}_3 &= 0.175 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{17.7 \text{ g KNO}_3} \end{aligned}$$

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