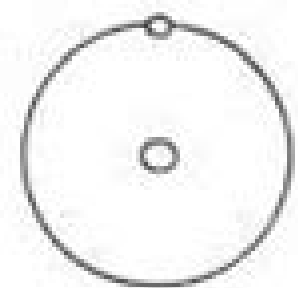


Name _____ Class _____ Date _____

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Coulomb's Law

1. The diagram is of a hydrogen atom.
 - a. Label the proton in the nucleus with a + sign and the orbital electron with a - sign.
 - b. The electrical interaction between the nucleus and the orbital electron is a force of _____ (attraction) (repulsion).
 - c. According to Coulomb's law,



$$F = k \frac{q_1 q_2}{r^2}$$

If the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be _____

(greater) (less)

and if the distance between the nucleus and electron were greater, the force would be _____ (greater) (less).

If the distance between the nucleus and electron were doubled, the force would be (1/4 as much) (1/2 as much) (two times as much) (4 times as much).

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