

Online Lab: Stoichiometry of Copper and Silver Nitrate Reaction

In this experiment, a coil of copper wire is reacted with silver nitrate solution. By measuring the mass of silver nitrate that dissolves in water to make the solution, and measuring the mass of copper that reacts with it, we can determine whether copper(I) nitrate or copper(II) nitrate is formed as the product along with silver metal. Depending on the product formed, the balanced chemical equation will be one of two possibilities:



Watch the video "Stoichiometry: Copper and Silver Nitrate Lab", posted by Seth Furlow on YouTube. Fill out the data table, then answer the questions below. (Show all of your work.)

<https://www.youtube.com/watch?v=rERbA7dFVFE>

Substance	Mass (g)
Beaker	50.7 g
AgNO ₃ solid	1.7049 g
Copper wire before reaction	3.7199 g
Copper wire after reaction	3.3897 g
Beaker and dried Ag	51.7 g

Copper Silver Nitrate Lab Answers

Peipei Pang



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