

Stoichiometry Lab: Iron and Copper Sulfate Sample data

Data and observations :

mass of copper (II) sulfate pentahydrate	<u>3.27</u>	g
mass of iron:	<u>0.50</u>	g
mass of filter paper	<u>1.28</u>	g
mass of filter paper and dry copper:	<u>1.84</u>	g
mass of dry copper	<u>0.56</u>	g

Observations of reaction:

When the iron was added a precipitate began to form. The precipitate was reddish in color. The solution remained a blue color but was somewhat lighter in color.

Calculations:

1&2. The single replacement equation will be: $\text{Fe} + \text{CuSO}_4 \longrightarrow \text{FeSO}_4 + \text{Cu}$

3. Fe and CuSO_4 are the reactants and FeSO_4 and Cu are the products.

4. The moles of copper (II) sulfate pentahydrate used were 0.013 mol.

$$3.27 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O} \times \frac{1 \text{ mol CuSO}_4 \cdot 5\text{H}_2\text{O}}{249.47 \text{ g CuSO}_4 \cdot 5\text{H}_2\text{O}} = 0.013 \text{ mol CuSO}_4 \cdot 5\text{H}_2\text{O}$$

5. The moles of iron used were 0.009 mol Fe.

$$0.50 \text{ g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.009 \text{ mol Fe}$$

6. The moles of copper produced were 0.009 mol Cu.

$$0.56 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.009 \text{ mol Cu}$$

Copper Sulfate And Iron Lab Answers

**David Martindill, Joanna
Haywood, Sheila Tarpey**



Copper Sulfate And Iron Lab Answers:

Experiment Station Record United States. Office of Experiment Stations,1942 Experiment Station Record U.S. Office of Experiment Stations,United States. Agricultural Research Service,United States. Office of Experiment Stations,1947

Laboratory Experiments in General Chemistry and Qualitative Analysis George Willard Watt,1951 *Annual Report - Maryland Agricultural Experiment Station* Maryland Agricultural Experiment Station,1894 **CHEMISTRY**

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United States. Office of Experiment Stations,1927 Information Circular
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Indexes to the Oak Ridge National Laboratory Master Analytical Manual Oak Ridge National Laboratory,1965
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