

# STATES OF MATTER

All matter is made of particles. The state of matter is determined by the speed of the particles and the strength of the attraction between them. Energy must be added or removed in order for something to change from one state to another.

## SOLID

Solids have definite shape and volume. The particles in a solid vibrate in place because they don't move fast enough to overcome the attraction between them.



## LIQUID

Liquids have definite volume but take the shape of the container holding them. The particles in a liquid move just fast enough to overcome some of the attraction between them and slide past one another.



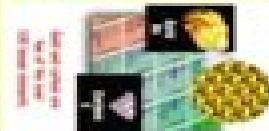
## GAS

Gases have no definite shape or volume. The particles in a gas move fast enough to overcome almost all attraction between them and move independently of one another. Gases can expand to fill the container they are in. Gases can be compressed.



## PLASMA

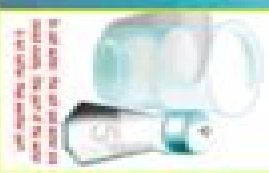
Plasma is extremely hot and highly electrically charged. It is the fourth state of matter.



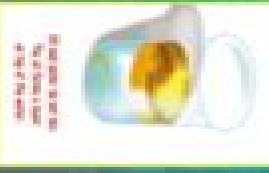
Plasma is a state of matter that exists at extremely high temperatures. It is composed of ions and free electrons.



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# PHYSICAL & CHEMICAL CHANGES

## PHYSICAL CHANGES

During a physical change, matter changes its appearance but remains chemically the same. The molecules that make up the matter are not rearranged.

### Examples of Physical Changes



## CHEMICAL CHANGES

During a chemical change, a chemical reaction takes place and one or more new substances are created. The molecules that make up the matter change when bonds break or form between the atoms that make up the molecules.

### Does That a Chemical Change Has Occurred



# ACIDS & BASES

## Acids



Acids are sour, conduct electricity, and react with metals to produce hydrogen gas. Lemon contains acid.

## Bases



Bases are bitter, feel slippery, and react with fats or oils to produce soap. Baking soda is a base.

## Neutral Substances



A neutral substance is not an acid or base. Pure water, for example, is a neutral substance.

## Indicators

Litmus paper is a chemical indicator. It changes color to indicate whether a solution is an acid, a base, or neutral. It does not tell how strong the acid or base is.

Indicator	Acidic Solution	Basic Solution	Neutral Solution
Litmus Paper	Red	Blue	Purple
Phenolphthalein	Colorless	Pink	Colorless
Bromothymol Blue	Yellow	Blue	Green

A pH scale indicates how acidic or basic a solution is. It is numbered from 0 to 14. Acids have a pH lower than 7. The lower the number, the more acidic the solution is. Bases have a pH greater than 7. The higher the number, the more basic a solution is. Substances with a pH of 7 are neutral.

A universal indicator changes color when it comes in contact with an acidic or basic solution. Comparing the new color to the color on a pH color chart indicates the pH level of the solution.



# MATTER

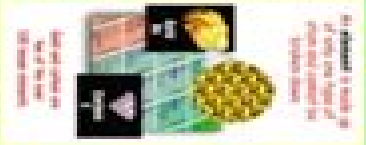
**Matter** is anything that takes up space and has mass. Matter can be categorized as a pure substance or a mixture.

A pure substance has the same composition throughout and pure substances often have specific physical and chemical properties.

A mixture forms when two or more substances are physically combined. The parts of a mixture can be physically separated as all the substances keep their original properties.

The parts of a homogeneous mixture are evenly distributed.

The parts of a heterogeneous mixture are not evenly distributed.



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the first five years of the award

**Subject Catalog, 1982** Library of Congress,1982

*The Publishers Weekly* ,1885

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    - Chapter 2: Essential Elements of Copyright C Mcdonald Publishing Co Basic Chemistry
    - Chapter 3: Copyright C Mcdonald Publishing Co Basic Chemistry in Everyday Life
    - Chapter 4: Copyright C Mcdonald Publishing Co Basic Chemistry in Specific Contexts
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- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Copyright C Mcdonald Publishing Co Basic Chemistry.

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