

# THE FOLLOWING QUESTIONS ARE PRACTICE QUESTIONS FOR THE CSN CHEMISTRY PLACEMENT EXAM

The chemistry placement test is used to assess your preparation for General Chemistry I, CHEM 121 at CSN. The test is a standardized test that will examine your knowledge of chemistry and basic mathematical skills. You will be provided scratch paper, a Periodic Table of the Elements and a nonprogrammable scientific calculator. The following sample questions are intended only as a guide to show you the relative level of difficulty and example topics covered in the Chemistry Placement Exam. The actual exam will have a time limit of 45 minutes to answer 44 multiple choice questions. Click the following link to view a [Periodic Table](#).

**MULTIPLE CHOICE.** Choose the single choice that best answers the question.

- 1) How many total atoms are in the formula



A) 12                      B) 14                      C) 9                      D) 8

- 2) How many electrons are in  $\text{Br}^-$ ?

A) 4                      B) 7                      C) 36                      D) 34

- 3) What are the coefficients for the following reaction when it is properly balanced?



A) 2, 3, 6, 1                      B) 6, 1, 3, 2  
C) 2, 3, 1, 6                      D) 2, 1, 1, 3

- 4) Given that  $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$ , when 4.50 mol of  $\text{H}_2\text{O}$  are formed, the amount of NO formed is

A) 3.00 mol                      B) 6.75 mol  
C) 1.50 mol                      D) 4.50 mol

- 5) What type of a reaction occurs when a sodium hydroxide solution is mixed with an acetic acid solution?

A) gas evolution  
B) oxidation-reduction  
C) acid-base neutralization  
D) precipitation

- 6) Which of the following atoms has the smallest radius?

A) Ne                      B) O  
C) Li                      D) B

- 7) What is the molecular geometry if you have a double bond, a single bond and 1 lone pair around the central atom?

A) trigonal pyramidal  
B) linear  
C) tetrahedral  
D) bent

- 8) What volume of 12.0 M HCl is required to make 75.0 mL of 3.50 M HCl?

A) 0.560 mL                      B) 257 mL  
C) 560. mL                      D) 21.9 mL

$$9) \frac{(2.80 \times 10^{-3})(1.28 \times 10^1)}{(3.20 \times 10^{-4})} = \underline{\hspace{2cm}}$$

A)  $1.27 \times 10^1$                       B)  $2.0 \times 10^2$   
C)  $1.2 \times 10^6$                       D)  $1.12 \times 10^2$

- 10) Of the molecules below, only \_\_\_\_\_ is polar.

A)  $\text{CH}_4$                       B)  $\text{SbF}_5$   
C)  $\text{CO}_2$                       D)  $\text{AsH}_3$

- 11) Which one of the following is a metalloid?

A) S                      B) Ge                      C) C                      D) Pb

- 12) A certain atom has a charge of -2 and has 16 protons and 16 neutrons in the nucleus. How many electrons does this atom possess?

A) 34                      B) 14                      C) 32                      D) 18

- 13) How many moles of pyridine ( $\text{C}_5\text{H}_5\text{N}$ ) are contained in 3.13 g of pyridine?

A) 0.00404                      B) 25.3  
C) 0.0396                      D) 0.319

- 14) Solve the following equation for x:

$$4x - b = 2x + c$$

A)  $c - b - 6$                       B)  $\frac{c + b}{2}$   
C)  $\frac{c + b}{6}$                       D)  $\frac{c - b}{2}$

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### **Table of Contents College Chemistry Placement Test Practice Uic**

1. Understanding the eBook College Chemistry Placement Test Practice Uic
  - The Rise of Digital Reading College Chemistry Placement Test Practice Uic
  - Advantages of eBooks Over Traditional Books
2. Identifying College Chemistry Placement Test Practice Uic
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an College Chemistry Placement Test Practice Uic
  - User-Friendly Interface
4. Exploring eBook Recommendations from College Chemistry Placement Test Practice Uic

- Personalized Recommendations
  - College Chemistry Placement Test Practice Uic User Reviews and Ratings
  - College Chemistry Placement Test Practice Uic and Bestseller Lists
5. Accessing College Chemistry Placement Test Practice Uic Free and Paid eBooks
    - College Chemistry Placement Test Practice Uic Public Domain eBooks
    - College Chemistry Placement Test Practice Uic eBook Subscription Services
    - College Chemistry Placement Test Practice Uic Budget-Friendly Options
  6. Navigating College Chemistry Placement Test Practice Uic eBook Formats
    - ePub, PDF, MOBI, and More
    - College Chemistry Placement Test Practice Uic Compatibility with Devices
    - College Chemistry Placement Test Practice Uic Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of College Chemistry Placement Test Practice Uic
    - Highlighting and Note-Taking College Chemistry Placement Test Practice Uic
    - Interactive Elements College Chemistry Placement Test Practice Uic
  8. Staying Engaged with College Chemistry Placement Test Practice Uic
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers College Chemistry Placement Test Practice Uic
  9. Balancing eBooks and Physical Books College Chemistry Placement Test Practice Uic
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection College Chemistry Placement Test Practice Uic
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine College Chemistry Placement Test Practice Uic
    - Setting Reading Goals College Chemistry Placement Test Practice Uic
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of College Chemistry Placement Test Practice Uic

- Fact-Checking eBook Content of College Chemistry Placement Test Practice Uic
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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