

DETERMINATION OF THE FORMULA OF COPPER (II) SULFATE

Purpose

To determine the number of moles of water in a mole of hydrated copper (II) sulfate is written as $\text{CuSO}_4 \cdot n \text{H}_2\text{O}$.

Introduction

Vocabulary: Salt, anhydrous salt, hydrated salt, crystal

A hydrate is a compound that contains water in its crystal structure. The water may be removed from the salt in the laboratory by heating the salt. The salt without the water is called an anhydrous salt. Here are some examples of hydrated salts:



In a hydrate the water molecules are a distinct part of the compound but are joined to the salt by connections that are weaker than the connections in the salt or the connections in the water molecules. Notice we use a dot to connect the water units to the salt formula.

In this experiment you will measure the mass of copper(II) sulfate, then remove the water from the crystals and measure the mass of the anhydrous salt. The data gathered will allow you to determine the percent water in the hydrated salt and also the empirical formula for the hydrated salt (McDougal Litel 2007).

Prelab question—see page 3

Materials

Bunsen burner, hydrated copper (II) sulfate, test tube, test tube holders

Procedure

1. Weigh a clean dry test tube.
2. Put about 1cm depth of copper (II) sulfate in the test tube and reweigh.
3. Record these weightings and those which follow in a table.
4. Now, use the Bunsen to heat the test tube gently for a minute and then strongly for about 3 minutes. Hold the test tube with a test tube holder. Allow the tube to cool and then reweigh it.
5. Heat the test tube again for about 2 minutes. Allow it to cool and then reweigh.
6. If the last two readings differ by more than 0.05g, repeat the heating and weighing again.

Data

Observations: Record all observations including appearance of the salt and what happens during heating.

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Table of Contents Compound Fourmla Lab Report Example

1. Understanding the eBook Compound Fourmla Lab Report Example
 - The Rise of Digital Reading Compound Fourmla Lab Report Example
 - Advantages of eBooks Over Traditional Books
2. Identifying Compound Fourmla Lab Report Example
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Compound Fourmla Lab Report Example
 - User-Friendly Interface
4. Exploring eBook Recommendations from Compound Fourmla Lab Report Example
 - Personalized Recommendations
 - Compound Fourmla Lab Report Example User Reviews and Ratings
 - Compound Fourmla Lab Report Example and Bestseller Lists
5. Accessing Compound Fourmla Lab Report Example Free and Paid eBooks
 - Compound Fourmla Lab Report Example Public Domain eBooks
 - Compound Fourmla Lab Report Example eBook Subscription Services
 - Compound Fourmla Lab Report Example Budget-Friendly Options
6. Navigating Compound Fourmla Lab Report Example eBook Formats

- ePub, PDF, MOBI, and More
- Compound Fourmla Lab Report Example Compatibility with Devices
- Compound Fourmla Lab Report Example Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Compound Fourmla Lab Report Example
 - Highlighting and Note-Taking Compound Fourmla Lab Report Example
 - Interactive Elements Compound Fourmla Lab Report Example
- 8. Staying Engaged with Compound Fourmla Lab Report Example
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Compound Fourmla Lab Report Example
- 9. Balancing eBooks and Physical Books Compound Fourmla Lab Report Example
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Compound Fourmla Lab Report Example
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Compound Fourmla Lab Report Example
 - Setting Reading Goals Compound Fourmla Lab Report Example
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Compound Fourmla Lab Report Example
 - Fact-Checking eBook Content of Compound Fourmla Lab Report Example
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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