

# Concept-Development Practice Page

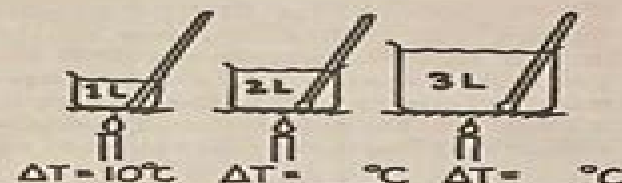
21-1

## Temperature and Heat

- Complete the table.

TEMPERATURE OF MELTING ICE	°C	32 °F	K
TEMPERATURE OF BOILING WATER	°C	212 °F	K

- Suppose you apply a flame and heat one liter of water, raising its temperature 10°C. If you transfer the same heat energy to two liters, how much will the temperature rise? For three liters? Record your answers on the blanks in the drawing at the right.



- A thermometer is in a container half-filled with 20°C water.

- When an equal volume of 20°C water is added, the temperature of the mixture is

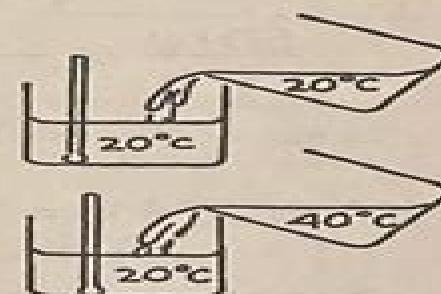
(10°C) (20°C) (40°C).

- When instead an equal volume of 40°C water is added, the temperature of the mixture will be

(20°C) (30°C) (40°C).

- When instead a small amount of 40°C water is added, the temperature of the mixture will be

(20°C) (between 20°C and 30°C) (30°C) (more than 30°C).



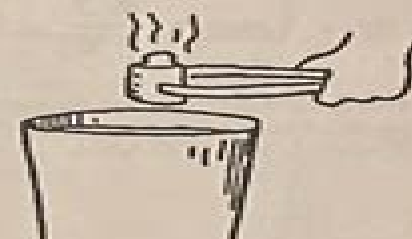
- A red-hot piece of iron is put into a bucket of cool water. Mark the following statements true (T) or false (F). (Ignore heat transfer to the bucket.)

- The decrease in iron temperature equals the increase in the water temperature. \_\_\_\_\_

- The quantity of heat lost by the iron equals the quantity of heat gained by the water. \_\_\_\_\_

- The iron and water both will reach the same temperature. \_\_\_\_\_

- The final temperature of the iron and water is halfway between the initial temperatures of each. \_\_\_\_\_



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**Mocktime Publication**



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