

IONIC & COVALENT BONDS

Name:

Date:

Period:

An _____ is the attraction between two oppositely charged ions. The chemical bond formed when two atoms share electrons is called a _____. Covalent bonds usually form between atoms of _____ and ionic bonds usually form when a _____ combines with a _____.

Activity A: How Do Ions Form?

Procedure:

1. Place three pairs of checkers (three red and three black) on your desk. The red represent electrons and the black represent protons.
2. Place nine pairs of checkers (nine red and nine black) in a separate group on your desk.
3. Move a red checker from the smaller group to the larger group.
4. Count the number of positive charges (protons) and negative charges (electrons) in each group.
5. Now sort the checkers into a group of four pairs and a group of eight pairs. Repeat Steps 3 and 4, this time moving two red checkers from the smaller group to the larger group.

Think it over:

1. What was the total charge on each group before you moved the red checkers (electrons)?
2. What was the charge on each group after you moved the checkers?
3. Based on this activity what do you think happens to the charge on an atom when it loses electrons?
4. Based on this activity what do you think happens to the charge on an atom when it gains electrons?

Activity B: Interpreting Data About Ions

Procedure:

1. Use the periodic table and the chart to the right to identify the charges of the ions in each ionic compound listed below.
2. Write the formulas for each compound.
 - a. sodium bromide;
 - b. lithium oxide;
 - c. magnesium sulfide;
 - d. aluminum fluoride;
 - e. potassium nitrate;
 - f. ammonium chloride

Ions and Their Charges		
Name	Charge	Symbol or Formula
Lithium	1+	Li^+
Sodium	1+	Na^+
Potassium	1+	K^+
Ammonium	1+	NH_4^+
Calcium	2+	Ca^{2+}
Magnesium	2+	Mg^{2+}
Aluminum	3+	Al^{3+}
Fluoride	1-	F^-
Chloride	1-	Cl^-
Iodide	1-	I^-
Bicarbonate	1-	HCO_3^-
Nitrate	1-	NO_3^-
Oxide	2-	O^{2-}
Sulfide	2-	S^{2-}
Carbonate	2-	CO_3^{2-}
Sulfate	2-	SO_4^{2-}
Phosphate	3-	PO_4^{3-}

Think it over:

1. How do you know how many of each ion to write in the formula?

Concept Review Chapter 6 Section Covalent Bonds

Answer Key

Mike Jess



Concept Review Chapter 6 Section Covalent Bonds Answer Key:

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Table of Contents Concept Review Chapter 6 Section Covalent Bonds Answer Key

1. Understanding the eBook Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - The Rise of Digital Reading Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Personalized Recommendations
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key User Reviews and Ratings

- Concept Review Chapter 6 Section Covalent Bonds Answer Key and Bestseller Lists
- 5. Accessing Concept Review Chapter 6 Section Covalent Bonds Answer Key Free and Paid eBooks
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key Public Domain eBooks
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key eBook Subscription Services
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key Budget-Friendly Options
- 6. Navigating Concept Review Chapter 6 Section Covalent Bonds Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key Compatibility with Devices
 - Concept Review Chapter 6 Section Covalent Bonds Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Highlighting and Note-Taking Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Interactive Elements Concept Review Chapter 6 Section Covalent Bonds Answer Key
- 8. Staying Engaged with Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concept Review Chapter 6 Section Covalent Bonds Answer Key
- 9. Balancing eBooks and Physical Books Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concept Review Chapter 6 Section Covalent Bonds Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Setting Reading Goals Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Fact-Checking eBook Content of Concept Review Chapter 6 Section Covalent Bonds Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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