

Part I: Preparation of ~ 0.1 M Stock Solution of NaCl

- Calculate how much NaCl you will need to prepare 100 mL of 0.1 M NaCl solution. The number of significant figures will be assigned based on the glassware you will use in the lab.
- Weigh the calculated amount using an analytical balance and a weighing boat.
- Record NaCl mass on your Data Sheet with the correct number of significant figures.
- Carefully transfer the solid into 100 mL volumetric flask making sure you are not losing any on the way. If any solid is stuck to the weighing boat, rinse it with DI water directly in to the flask using a squeeze bottle.
- Add DI water to the flask using a squeeze bottle until it is about 2/3 full.
- Put a stopper and gently swirl the flask to dissolve the solid at the bottom.
- Add more water until you are about 1/2 of the flask's neck distance from the marking.
- Take a Pasteur pipette equipped with a pipette bulb and add water dropwise until you reach the mark. Remember, that the lowest meniscus line should be lying on the flask marking.
- Put the stopper and swirl your flask again to make sure the solution is well mixed.

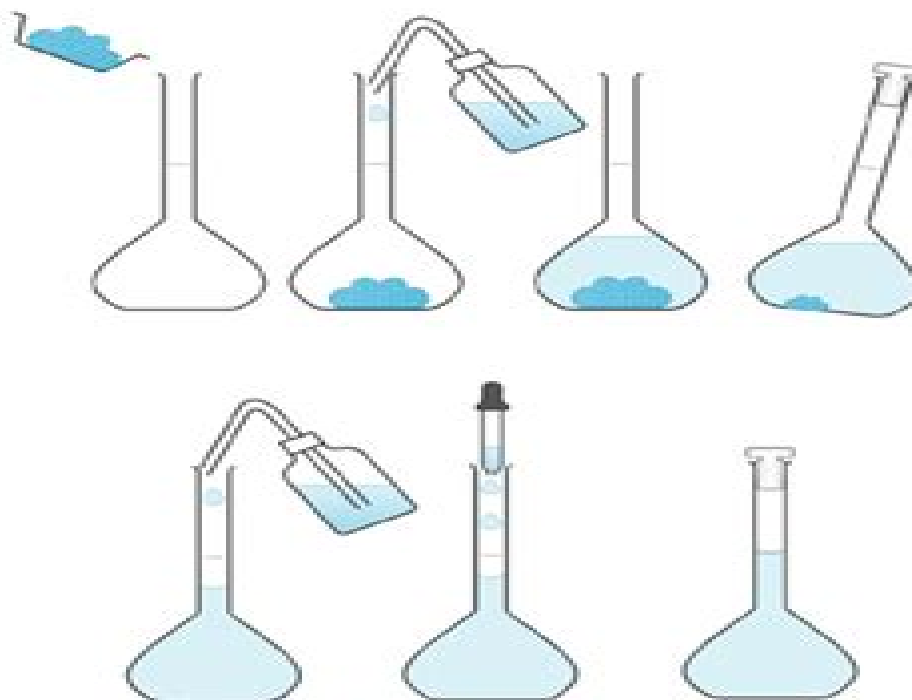


Figure 1 - How to prepare a solution of a given volume using a volumetric flask

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