Configuration Writing Practice

- A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.
- No: 1s², 2s², 2p⁶, 3s¹
- Pb: 1s² 2s²2p° 3s²3p° 4s² 3d¹0 4p° 5s² 4d¹0 5p° 6s² 4f¹0
- 3. Sr: 1s2 2s22p6 3s23p6 4s2 3d10 4p6 5s2
- U: 1s²,2s²,2p⁶,3s²,3p⁶,4s²,3d¹⁰,4p⁶,5s²,4d¹⁰5p⁶6s²4f¹⁴5d¹⁰6p⁶7s²5f⁶
- 5. N: 1s2, 2s2, 2p3
- B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: N³ is 1s² 2s² 2p⁶. It has three extra electrons

B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: N3-is 1s2 2s2 2p6. It has three extra electrons

- O²: 1s²,2s²,2p⁶
- Fe²⁺: 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d⁵
- 13. B^{3*}: 1s²
- Ni²⁺ 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d⁶
- K': 1s² 2s²2p⁶ 3s²3p⁶
- Co³⁻ 1s² 2s³2p⁶ 3s³3p⁶ 4s² 3d⁵

C. For the following electron configurations determine the possible elements (or ions) they may represent

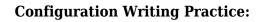
- 17. 1s2 2s22p6 3s23p6 4s2 3d10 4p4 : Se
- 18. 1s2 2s22p6 3s23p6 4s2 3d10 4p6 : Kr
- 19. 1s2 2s22p6 3s23p6 4s2 3d10 4p6 5s2 4d10 5p6 6s2 4f10 : Dy
- 1s2 2s22p6 3s23p6 4s2 3d5 : Mn
- [Kr] 5s2 4d10 5p3: As
- 22. [Ar] 4s1 : K
- [Xe] 6s2 4f10 : Dy
- [Ne] 3s2 3p1 : Al

C. In the space below, write the full (unabbreviated) electron configurations of the following elements:

- sodium 1s² 2s²2p⁶ 3s¹
- iron 1s² 2s²2p° 3s²3p° 4s² 3d°
- bromine
- 4) barium 1s² 2s²2p6 3s²3p6 4s² 3d²0 4p6
- 5) neptunium 1s²,2s²,2p⁶,3s²,3p⁶,4s²,3d¹⁰,4p⁶,5s²,4d¹⁰5p⁶6s²4f¹⁴5d¹⁰6p⁶7s²5f⁶

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