

Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2, 2s^2, 2p^6, 3s^1$
2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{10}$
3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$
4. U: $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^6$
5. N: $1s^2, 2s^2, 2p^3$

B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: N^{3-} is $1s^2 2s^2 2p^6$. It has three extra electrons

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11. O^{2-} : $1s^2, 2s^2, 2p^6$
12. Fe^{2+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
13. B^{3+} : $1s^2$
14. Ni^{2+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$
15. K^+ : $1s^2 2s^2 2p^6 3s^2 3p^6$
16. Co^{3+} : $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$

C. For the following electron configurations determine the possible elements (or ions) they may represent

17. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$: Se
18. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$: Kr
19. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{10}$: Dy
20. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$: Mn
21. [Kr] $5s^2 4d^{10} 5p^3$: As
22. [Ar] $4s^1$: K
23. [Xe] $6s^2 4f^{10}$: Dy
24. [Ne] $3s^2 3p^1$: Al

C. In the space below, write the full (unabbreviated) electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$
- 5) neptunium $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^6$

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Table of Contents Configuration Writing Practice

1. Understanding the eBook Configuration Writing Practice
 - The Rise of Digital Reading Configuration Writing Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Configuration Writing Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Configuration Writing Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Configuration Writing Practice
 - Personalized Recommendations
 - Configuration Writing Practice User Reviews and Ratings
 - Configuration Writing Practice and Bestseller Lists
5. Accessing Configuration Writing Practice Free and Paid eBooks

- Configuration Writing Practice Public Domain eBooks
- Configuration Writing Practice eBook Subscription Services
- Configuration Writing Practice Budget-Friendly Options
- 6. Navigating Configuration Writing Practice eBook Formats
 - ePub, PDF, MOBI, and More
 - Configuration Writing Practice Compatibility with Devices
 - Configuration Writing Practice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Configuration Writing Practice
 - Highlighting and Note-Taking Configuration Writing Practice
 - Interactive Elements Configuration Writing Practice
- 8. Staying Engaged with Configuration Writing Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Configuration Writing Practice
- 9. Balancing eBooks and Physical Books Configuration Writing Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Configuration Writing Practice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Configuration Writing Practice
 - Setting Reading Goals Configuration Writing Practice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Configuration Writing Practice
 - Fact-Checking eBook Content of Configuration Writing Practice
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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