

10**CHEMICAL QUANTITIES****Chapter Test B****A. Matching**

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

Column A

- _____ 1. the percent by mass of each element in a compound
- _____ 2. 0°C and 101.3 kPa or 1 atmosphere
- _____ 3. the lowest whole-number ratio of the atoms of elements in a compound
- _____ 4. the species present in a substance—usually atoms, molecules, or formula units
- _____ 5. 6.02×10^{23} representative particles of a substance
- _____ 6. the mass of an atom of an element
- _____ 7. the mass of one mole of any element or compound
- _____ 8. 22.4 L of any gas measured at STP

Column B

- a. molar mass
- b. empirical formula
- c. atomic mass
- d. molar volume
- e. representative particle
- f. percent composition
- g. one mole
- h. standard temperature and pressure

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 9. How many oxygen atoms are in 10 formula units of $\text{Al}_2(\text{SO}_4)_3$?
 - a. 10 atoms O
 - b. 7 atoms O
 - c. 70 atoms O
 - d. 120 atoms O
- _____ 10. How many molecules are in 4.50 moles of H_2O ?
 - a. 450 molecules
 - b. 2.71×10^{28} molecules
 - c. 6.02×10^{23} molecules
 - d. 3.00 molecules
- _____ 11. How many moles are in 8.5×10^{25} molecules of CO_2 ?
 - a. 1.4×10^2 mol
 - b. 7.1×10^{-25} mol
 - c. 5.1×10^{19} mol
 - d. 8.5×10^{25} mol

Core Teaching Resources Chapter 8

Thomas Griffiths



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