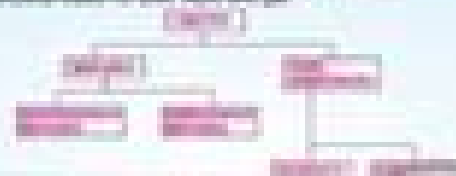


1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various sub-classifications of matter are given below:
- Matter is classified into pure and impure substances.
- It is classified into elements and compounds.
- It is classified into metals and non-metals.
- It is classified into pure and impure substances.
- It is classified into crystalline and amorphous substances.
- It is classified into ideal and real gases.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are physical and chemical properties.
 - Physical properties are those which can be measured without changing the chemical composition of the substance.
 - Chemical properties are those which can be measured by changing the chemical composition of the substance.
 - Colour, odour, melting point, density, etc. are some physical properties.
 - Combustion, corrosion, etc. are some chemical properties.
 - Quantitative measurement of physical properties is called measurement of matter.
 - Properties of matter are measured by using standard units.
 - Some important quantities and their units are:
- | Quantity | SI Unit |
|---------------------------|-------------------------------|
| Mass (g) | Kilogram (kg) |
| Volume (cm ³) | Cubic metre (m ³) |
| Temperature (°C) | Kelvin (K) |

3. CONVERSION OF UNITS AND DIMENSIONS

- 1 kg = 1000 g, 1 m = 1000 mm, 1 L = 1000 mL, 1 dm = 10 cm, 1 cm = 10 mm, 1 m = 1000 mm, 1 km = 1000 m, 1 h = 60 min, 1 min = 60 s, 1 day = 24 h, 1 year = 365 days.

4. SIGNIFICANT FIGURES

- 1.1 + 0.0004 = 1.1004, 1.1004 + 0.0001 = 1.1005, 1.1005 + 0.0001 = 1.1006, 1.1006 + 0.0001 = 1.1007, 1.1007 + 0.0001 = 1.1008.

1. UNCERTAINTY IN MEASUREMENT

- Measurement is a process of determining the value of a physical quantity.
- Measurement is always uncertain.
- Measurement is always uncertain.

- Significant figures are those digits which are known with certainty plus one additional uncertain digit.

2. RULES FOR DETERMINING THE NUMBER OF SIGNIFICANT FIGURES

- Non-zero digits are significant.
- Zeros between two non-zero digits are significant.
- Zeros at the end or right of a number are significant if they are on the right side of a decimal point.
- Leading zeros are not significant.
- In a number written in scientific notation, all digits are significant.
- Exact values are the values of certain measurements for which there is no uncertainty.
- When adding or subtracting, the result should have the same number of significant figures as the least precise measurement.
- When multiplying or dividing, the result should have the same number of significant figures as the least precise measurement.
- In a measurement, the number of significant figures is the number of digits which are known with certainty plus one additional uncertain digit.

Aakash Chemistry Lab Manual

Clemens Wendtner



Aakash Chemistry Lab Manual:

Quantitative Analysis Laboratory Manual Dr. Bupinderpal Singh ,Dr. Bikramjit Singh,Mr. Akash Bans,2025-06-29

Energetic Materials Mark J. Mezger,Kay J. Tindle,Michelle Pantoya,Lori J. Groven,Dilhan Kalyon,2017-07-06 This book will take an in depth look at the technologies processes and capabilities to develop and produce next generation energetic materials for both commercial and defense applications including military mining operations oil production and well perforation and construction demolition It will serve to highlight the critical technologies latest developments and the current capability gaps that serve as barriers to military fielding or transition to the commercial marketplace It will also explain how the processing technologies can be spun out for use in other non energetics related industries Textile

Chemist and Colorist ,1995 **Who's Who in Science and Engineering 2008-2009** Who's Who Marquis,Marquis Who's

Who,2007-12 **Books in Print Supplement** ,2002 Laboratory Manual for Chemistry ,1966 Who's who in the West

,1998 Elementary Practical Chemistry George Samuel Newth,1906 **Practical/Laboratory Manual Chemistry Class**

- **XI** Er. Meera Goyal,2021-05-29 1 Basic Laboratory Techniques 1 To cut a glass tube or glass rod 2 To bend the glass rod at an angle 3 To draw a glass jet from a glass tube 4 To bore a cork and fit a glass tube into it Viva Voce 2 Characterisation and Purification of Chemical Substances 1 To determine the melting point of the given unknown organic compound and its identification simple laboratory technique Viva Voce 2 To determine the boiling point of a given liquid when available in small quantity simple laboratory method Viva Voce 3 To prepare crystals of pure potash alum $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$ from the given impure sample 4 To prepare the pure crystals of copper sulphate from the given crude sample 5 To prepare pure crystals of benzoic acid from a given impure sample Viva Voce 3 Measurement of pH Values 1 To determine the pH value of vegetable juices fruit juices tap water and washing soda by using universal pH paper 2 To determine and compare the pH values of solutions of strong acid HCl and weak acid CH_3COOH of same concentration 3 To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper 4 To study the pH change by common ion CH_3COO^- ion in case of weak acid CH_3COOH 5 To determine the change in pH value of weak base NH_4OH in presence of a common ion NH_4^+ Viva Voce 4 Chemical Equilibrium 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions 2 To study the shift in equilibrium between $Co(H_2O)_6^{2+}$ and Cl^- ions by changing the concentrations of either of the ions Viva Voce 5 Quantitative Analysis 1 To prepare M 10 oxalic acid solution by direct weighing method 2 To prepare M 10 solution of sodium carbonate by direct weighing method 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution Viva Voce 6 Qualitative Analysis Analysis of Anions Analysis of Cations Viva Voce 7 Detection of Elements in Organic Compounds 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne's test 2 To

detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne's test Viva Voce

INVESTIGATORY PROJECTS

1 Checking of Bacterial Contamination in Water

1 To check the bacterial contamination in drinking water by testing sulphide ions Viva Voce

2 Methods of Water Purification

1 To purify water from suspended impurities by using sedimentation

2 To purify water by boiling

3 To purify water by distillation method

4 To purify water by reverse osmosis technique

5 To purify water by GAC method

6 To purify water by bleach treatment

7 To purify water by oxidising agent

8 To purify water by ozone treatment method Viva Voce

3 Water Analysis

1 To test the hardness of different water samples Viva Voce

4 Foaming Capacity of Various Soaps

1 To compare the foaming capacity of different washing soaps

2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap Viva Voce

5 Tea Analysis

1 To study the acidity of different samples of tea leaves tea by using pH paper Viva Voce

6 Analysis of Fruits and Vegetable Juices

1 To analysis the fruit and vegetable juices for the constituent present in them Viva Voce

7 Rate of Evaporation

1 To study the rate of evaporation of different liquids IViva Voce

8 Effect of Acids and Bases on Tensile Strength of Fibres

1 To compare the tensile strength of natural fibres and synthetic fibres

2 To study the effect of acids and bases on tensile strength of different fibres Viva Voce

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