

\* Show that, the entropy remains constant during a reversible cyclic change.

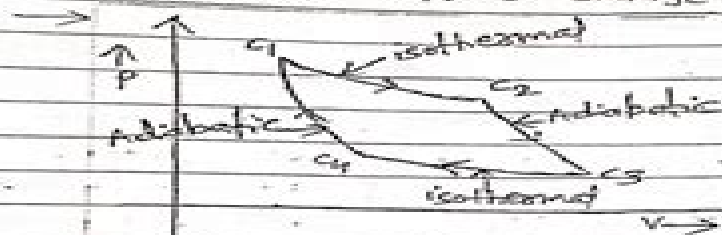


Fig: Carnot's cycle

Consider a reversible cycle  $C_1 C_2 C_3 C_4 C_1$  as shown in fig. It consists of 4 operations  
 1) From  $C_1$  to  $C_2$  heat  $Q_1$  is absorbed by the system at temp.  $T_1$ . The change in entropy of the system is  $S_1$ . The change is positive.

2) From  $C_2$  to  $C_3$ , no heat is absorbed or rejected by the system, hence no change in entropy

3) From  $C_3$  to  $C_4$ , heat  $Q_2$  is rejected by the system at temp.  $T_2$ . The change in entropy is  $S_2$ . The change is negative.

4) From  $C_4$  to  $C_1$ , no heat is absorbed or rejected by the system, hence no change in entropy.

$\therefore$  Total change in entropy of a working substance during reversible cycle is  $\frac{Q_1}{T_1} = \frac{Q_2}{T_2} \Rightarrow \frac{Q_1}{T_1} - \frac{Q_2}{T_2} = 0$

change in entropy of the system of source & sink is zero. Thus during reversible cycle, the entropy of the universe remains constant. i.e.  $\Delta S_R = 0$ .

\* state & explain the principle of increase of entropy or explain entropy of universe in irreversible processes.  
 → Efficiency of Carnot's cycle is

$$\eta = 1 - \frac{Q_2}{Q_1}$$

But, this efficiency is less than that of a reversible engine working between the same two temps  $T_1$  &  $T_2$ . for which, efficiency

$$\eta = 1 - \frac{T_2}{T_1}$$

$$\therefore 1 - \frac{Q_2}{Q_1} < 1 - \frac{T_2}{T_1}$$

# Applied Thermodynamics Notes

**M. R. Riazi**



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