

Balancing a Chemical Equation – Rules

1. Count your atoms and set up your T-Chart.
2. You can only change coefficients, NEVER change the subscripts.
3. Balance Hydrogen last.
4. Balance Oxygen second-to-last.

Balancing a Chemical Equation – Worked Example



Note – this is a difficult problem. Most require less steps

Step 1: Count your atoms and set up your T-Chart.
Result: Nitrogen and Oxygen are balanced but Hydrogen is not balanced in the products side.



T-Chart Step 1		
Reactants	Atoms	Products
1	Nitrogen	1
2	Oxygen	2 + 1 = 3
3	Hydrogen	2

Step 2: Add more hydrogen to the products by putting a "2" coefficient in front of H_2O .
Result: the two most hydrogen in the products, two more Oxygen and Hydrogen are unbalanced, set them aside for the time being.



T-Chart Step 2		
Reactants	Atoms	Products
2	Nitrogen	1
2	Oxygen	2 + 2 + 2 = 6
6	Hydrogen	4 + 2 = 6

Step 3: Add more Oxygen and Oxygen to the reactants by putting a "3" coefficient in front of O_2 and O_2 .
Result: the two most Oxygen and Hydrogen in the reactants but now nitrogen is unbalanced, set aside Nitrogen first.



T-Chart Step 3		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	1
6 + 6 = 12	Oxygen	2 + 2 + 2 = 6
6 + 2 = 8	Hydrogen	4 + 2 = 6

Step 4: Add more Nitrogen to the reactants by putting a "2" coefficient in front of NO .
Result: Nitrogen and Oxygen are balanced but now Hydrogen is not balanced.



T-Chart Step 4		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2 + 2 = 4
6 + 6 = 12	Oxygen	2 + 2 + 2 + 2 = 8
6 + 2 = 8	Hydrogen	4 + 2 = 6

Step 5: Add more Hydrogen to the reactants by putting a "3" coefficient in front of NH_3 .
Result: 3 balanced. In balanced now all reactants Oxygen are the products side.



T-Chart Step 5		
Reactants	Atoms	Products
3 + 3 = 6	Nitrogen	2 + 2 = 4
6 + 6 = 12	Oxygen	2 + 2 + 2 + 3 = 9
9 + 3 = 12	Hydrogen	4 + 2 = 6

Balancing Chemical Equations Cheat Sheet



Rod Powers

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