Balancing Chemical Equations - Worksheet #2 Classifying Chemical Reactions

Most chemical reactions can be classified as one of these six basic reaction types:

- Synthesis (combination)
- · Decomposition
- Single Replacement (oxidation-reduction)
- · Double Replacement (precipitation)
- Acid-Base (neutralization)
- Combustion

Write balanced equations for each of the following examples:

Synthesis Reaction: A + B → AB

Sodium (s) + chlorine (g) \rightarrow sodium chloride (s)

2. Decomposition Reaction: AB → A + B

sodium bicarbonate (s) -> carbon dioxide (g) * sodium carbonate (s) * water (l)

3. Single Replacement Reactions: AB + C → A + CB

hydrochloric acid, HCl (aq) + magnesium (s) → magnesium chloride (aq) + hydrogen (g)

Double Replacement Reactions: AB + CD → AD + CB

calcium chloride (aq) + ammonium hydroxide (aq) -> calcium hydroxide (s) + ammonium chloride (aq)

Acid-Base Reactions: HA + BOH → BA + H₂O

Sulfuric Acid, H₂SO₄ (aq) + potassium hydroxide (aq) → potassium sulfate (aq) + water (I)

Combustion Reactions: Fuel + O₂(g) → CO₂(g) + H₂O(g)

Hexane, C₆H₁₄ (l) + oxygen (g) → carbon dioxide (g) + water (g)

Balancing Chemical Equations Study Guide Answers

H. Stephen Stoker, Danny V. White

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